



SC12 Reversible reactions and equilibria

Personal learning checklist | 1 of 1

GCSE Edexcel Separate Sciences | Chemistry Paper 1

Name _____ Class _____ Date _____

Tick one box for each statement. Red: need help. Amber: partly confident.
Green: can do independently.

I can	Red	Amber	Green
Reversibility and equilibrium			
Describe reversible reactions, their symbol and ammonium chloride decomposition and recombination.	☐	☐	☐
Explain dynamic equilibrium in a closed system using equal rates and constant, potentially unequal concentrations.	☐	☐	☐
Interpret rate-time and concentration-time graphs and distinguish equilibrium position from reaction completion.	☐	☐	☐
Changing conditions (Higher)			
Predict equilibrium shifts after concentration changes, including adding reactants or removing products (Higher).	☐	☐	☐
Predict temperature effects using the energy change of the written reaction (Higher).	☐	☐	☐
Predict pressure effects from gaseous coefficients, including equal gas totals (Higher).	☐	☐	☐
Explain why catalysts speed equilibrium attainment without changing equilibrium yield (Higher).	☐	☐	☐
Ammonia manufacture			
Describe Haber reactants, their sources, balanced reaction, temperature, pressure and iron catalyst.	☐	☐	☐
Explain Haber temperature and pressure compromises between yield, rate, energy and equipment costs (Higher).	☐	☐	☐
Describe ammonia removal and unreacted gas recycling and explain improved feedstock use.	☐	☐	☐

My next revision priority _____